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The following table shows the results of the experiment conducted on the effect of temperature on the rate of reaction between potassium dichromate and potassium iodide in the presence of hydrochloric acid.

The results show that the rate of reaction increases with an increase in temperature. This is because the molecules have more kinetic energy and collide more frequently and with more energy.

The rate of reaction is directly proportional to the temperature. This is shown by the graph of rate of reaction versus temperature, which is a straight line passing through the origin.

The following table shows the results of the experiment conducted on the effect of concentration on the rate of reaction between potassium dichromate and potassium iodide in the presence of hydrochloric acid.

The results show that the rate of reaction increases with an increase in concentration.

The rate of reaction is directly proportional to the concentration. This is shown by the graph of rate of reaction versus concentration, which is a straight line passing through the origin.

The following table shows the results of the experiment conducted on the effect of surface area on the rate of reaction between potassium dichromate and potassium iodide in the presence of hydrochloric acid.

The results show that the rate of reaction increases with an increase in surface area.

The rate of reaction is directly proportional to the surface area. This is shown by the graph of rate of reaction versus surface area, which is a straight line passing through the origin.

The following table shows the results of the experiment conducted on the effect of catalyst on the rate of reaction between potassium dichromate and potassium iodide in the presence of hydrochloric acid.